

Physics 492: Quantum Mechanics II

Lecture 10:

Hydrogenic Energy Spectrum
and Eigenfunctions (continued)

Last lecture we solved the T.I.S.E. for the bound states of hydrogen

Relative motion:
$$\hat{H}_e = \frac{\hat{p}_r^2}{2m_e} + \frac{\hbar^2 l(l+1)}{2mr^2} - \frac{e^2}{r}$$

$$\hat{H} \Psi_{n_r, l, m} = E_{n, l} \Psi_{n_r, l, m}$$

$$E_{n, l} = -\frac{R}{(n_r + l + 1)^2} = -\frac{R}{n^2}$$

$$R = 13.6 \text{ eV}, \quad n \equiv n_r + l + 1 = 1, 2, 3, \dots$$

$$\Psi_{n_r, l, m} = R_{n_r, l}(r) Y_l^m(\theta, \phi)$$

Radial wave function $R_{n_r, l}(r) = \frac{U_{n_r, l}(r)}{r}$
In dimensionless units $\bar{r} \equiv r/a_0$ $\bar{E}_b = E_b/E_0$

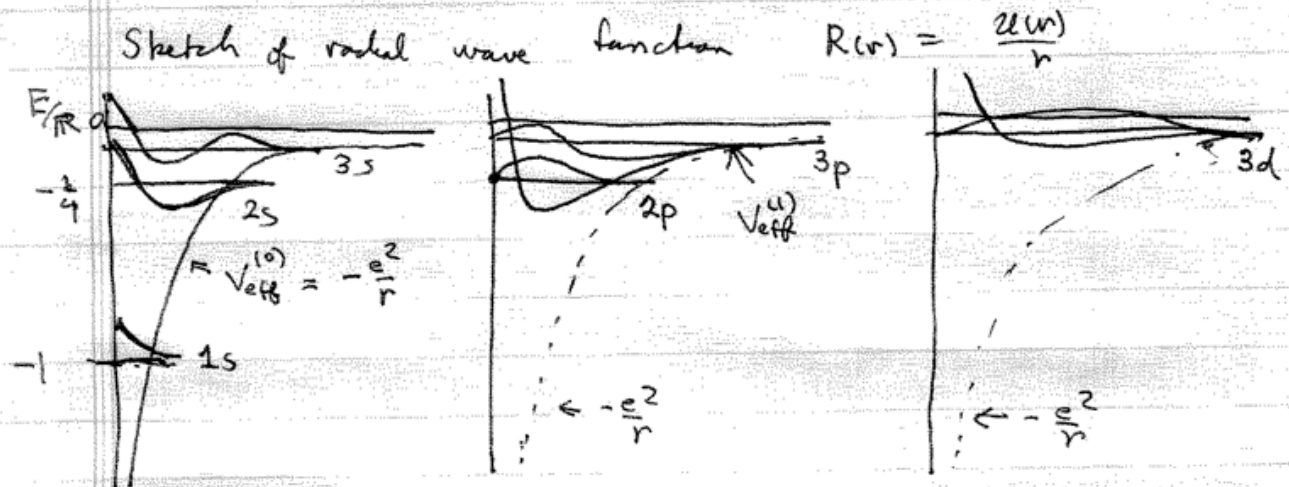
$$U_{n_r, l}(\bar{r}) = \bar{r}^{l+1} e^{-\bar{r}/n} F_{n_r, l}(2\bar{r}/n)$$

$$\bar{r} \equiv \sqrt{2|E_n|} = \frac{1}{n}$$

$$F_{n_r, l}(2\bar{r}/n) = N_{n_r, l} \underset{\text{norm}}{L_{n_r}^{2l+1}}(2\bar{r}/n) \quad (\text{associated Laguerre})$$

Laguerre polynomial, order p , $L_p(x) = e^x \frac{d^p}{dx^p} (x^p e^{-x})$

Associate Laguerre: $L_p^q(x) = (-1)^q \frac{d^q}{dx^q} L_{p+q}(x)$



Spectroscopic Notation	Several Normalized Time-Independent Eigenstates of Hydrogen
1S	$\varphi_{100} = \frac{2}{a_0^{3/2}} e^{-r/a_0} Y_0^0(\theta, \phi)$
2S	$\varphi_{200} = \frac{2}{(2a_0)^{3/2}} (1 - r/2a_0) e^{-r/2a_0} Y_0^0(\theta, \phi)$
2P	$\begin{pmatrix} \varphi_{211} \\ \varphi_{210} \\ \varphi_{21-1} \end{pmatrix} = \frac{1}{\sqrt{3}(2a_0)^{3/2}} \frac{r}{a_0} e^{-r/2a_0} \begin{pmatrix} Y_1^1(\theta, \phi) \\ Y_1^0(\theta, \phi) \\ Y_1^{-1}(\theta, \phi) \end{pmatrix}$
3S	$\varphi_{300} = \frac{2}{3(3a_0)^{3/2}} [3 - 2r/a_0 + 2(r/3a_0)^2] e^{-r/3a_0} Y_0^0(\theta, \phi)$
3P	$\begin{pmatrix} \varphi_{311} \\ \varphi_{310} \\ \varphi_{31-1} \end{pmatrix} = \frac{4\sqrt{2}}{9(3a_0)^{3/2}} \frac{r}{a_0} (1 - r/6a_0) e^{-r/3a_0} \begin{pmatrix} Y_1^1(\theta, \phi) \\ Y_1^0(\theta, \phi) \\ Y_1^{-1}(\theta, \phi) \end{pmatrix}$
3D	$\begin{pmatrix} \varphi_{322} \\ \varphi_{321} \\ \varphi_{320} \\ \varphi_{32-1} \\ \varphi_{32-2} \end{pmatrix} = \frac{2\sqrt{2}}{27\sqrt{5}(3a_0)^{3/2}} \left(\frac{r}{a_0}\right)^2 e^{-r/3a_0} \begin{pmatrix} Y_2^2(\theta, \phi) \\ Y_2^1(\theta, \phi) \\ Y_2^0(\theta, \phi) \\ Y_2^{-1}(\theta, \phi) \\ Y_2^{-2}(\theta, \phi) \end{pmatrix}$

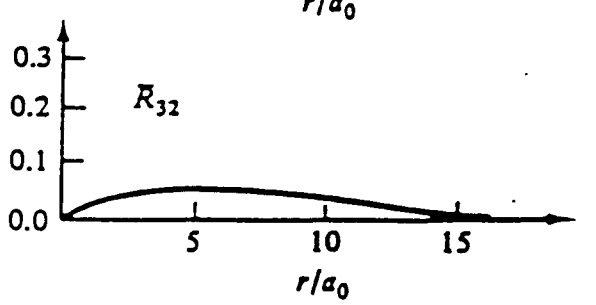
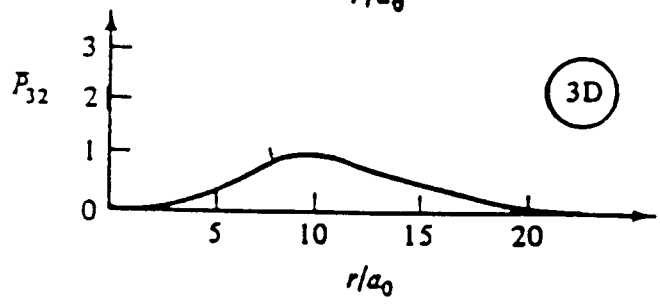
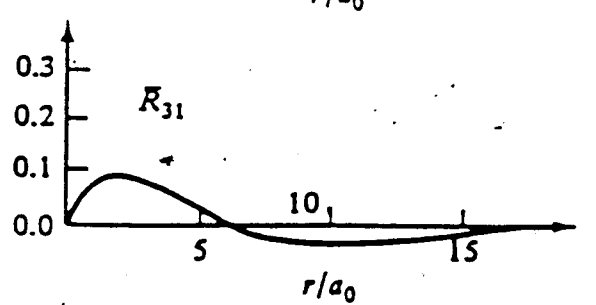
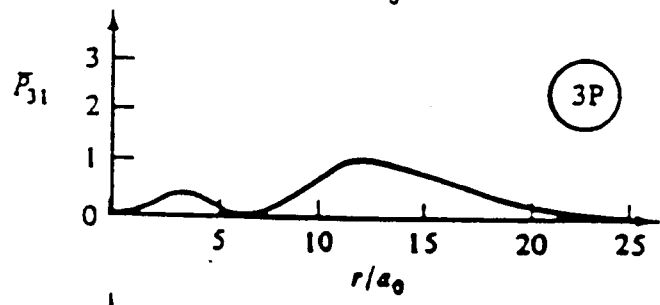
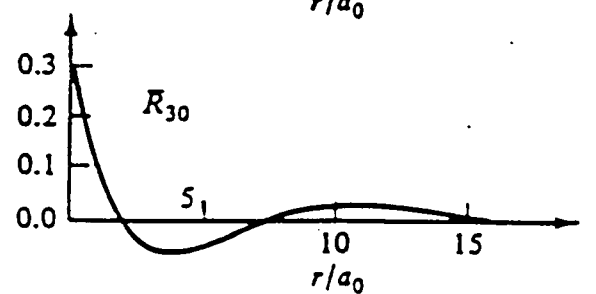
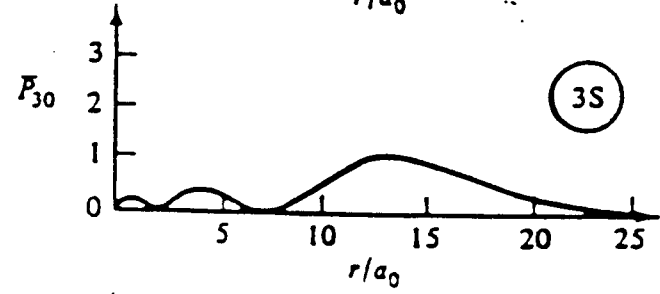
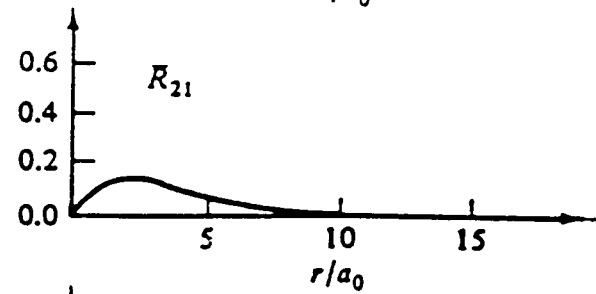
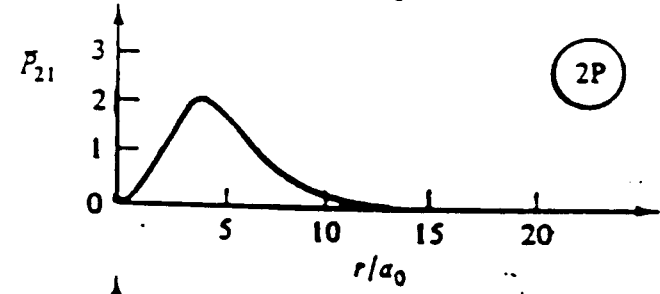
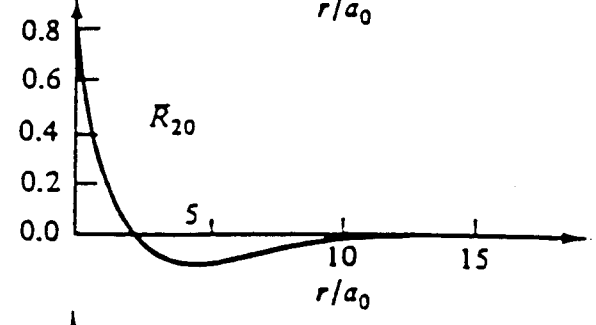
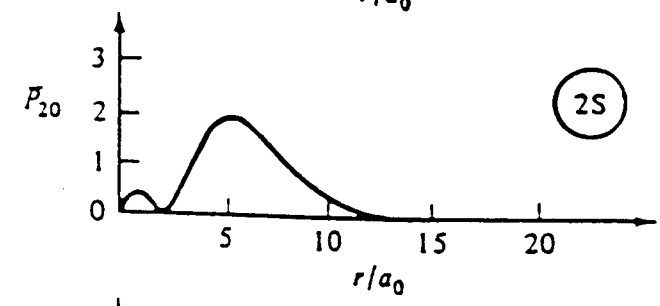
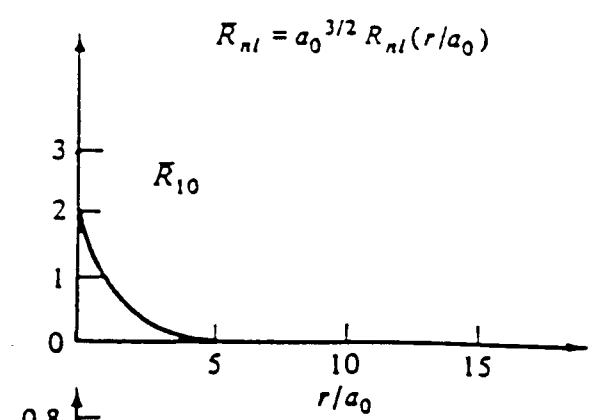
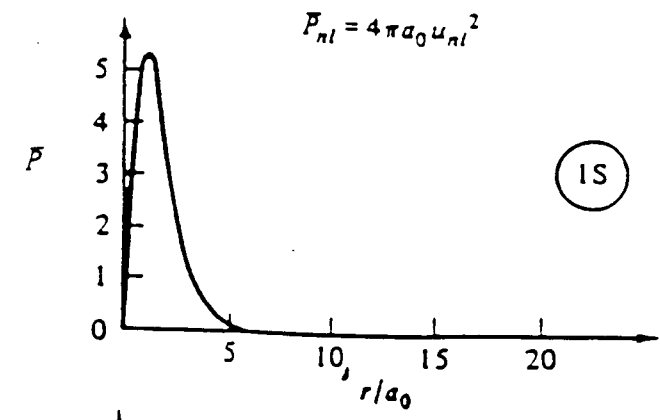


FIGURE 12-5.

FIG. 12-4

Normalizations: $\int d^3r |\psi_{n_r, l, m}|^2 = 1$

$$\Rightarrow \int dr \underbrace{r^2 |R_{n_r, l}(r)|^2}_{P(r)} \underbrace{\int d\Omega |Y_l^m|^2}_1 = 1$$

$P(r)$ = probability density to find electron in radius $r \rightarrow r dr$

$$P(r) = r^2 |R_{n_r, l}(r)|^2 = |u_{n_r, l}(r)|^2$$

units = $\frac{1}{(\text{Length})}$

Dimensionless connection: $r = a_0 \bar{r}$

$$\int dr r^2 |R(r)|^2 = \int dr |u(r)|^2$$

Define $u(r) \equiv \frac{1}{\sqrt{a_0}} \bar{u}(\bar{r} = \frac{r}{a_0})$

Putting it all together

$$u_{n_r, l}(r) = N_{n_r, l} \left(\frac{2r}{na_0}\right)^{l+1} e^{-\frac{r}{na_0}} L_{n_r}^{2l+1}\left(\frac{2r}{n a_0}\right)$$

$$N_{n_r, l} = \left(\frac{2}{na_0}\right)^{3/2} \sqrt{\frac{n_r!}{2n [(n+l)!]^3}}, \quad n = n_r + l + 1$$

"Circular" orbits: No radial excitation $n_r = 0$
 $\Rightarrow n = l + 1$, $l = n - 1 = l_{\max}$

Radial probability density

$$P(r) = r^2 |R_{n, l_{\max}}|^2 = |U_{0, n-1}(r)|^2$$

$$\propto |r^n e^{-\frac{r}{a_0 n}} \underbrace{L_0^{n-1}\left(\frac{2r}{a_0 n}\right)}_{\text{constant, } 0^{\text{th}} \text{ order poly}}|^2$$

$$\Rightarrow P(r) \propto r^{2n} e^{-2r/a_0 n}$$

Peak probability: $\frac{dP}{dr} = (2n r^{2n-1} - \frac{2}{a_0 n} r^{2n}) e^{-\frac{2r}{a_0 n}} = 0$

$$\Rightarrow \boxed{r_{\text{peak}} = n^2 a_0}$$

Expectation value $\langle r \rangle_{n, l, m} = \int d^3x r |\psi_{n, l, m}(\mathbf{x})|^2$

$$= \int_0^{\infty} dr r^3 |R_{n, l}|^2 = \int_0^{\infty} dr r |u_{n, l}|^2$$

$$\Rightarrow \langle r \rangle_{n, l} = a_0 \int_0^{\infty} d\bar{r} \bar{r} |\bar{u}_{n, l}(\bar{r})|^2 \quad (\text{independent of } m)$$

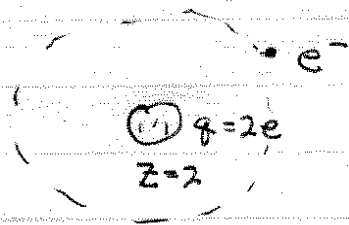
Example: Ground state $\bar{u}_{1s}(\bar{r}) = 2\bar{r} e^{-\bar{r}}$

$$\Rightarrow \langle r \rangle_{1s} = a_0 \left[\int_0^{\infty} d\bar{r} \bar{r}^3 e^{-2\bar{r}} \right] = \frac{3a_0}{2}$$

Generally $\boxed{\langle r \rangle_{n, l} = a_0 \left[\frac{3n^2 - l(l+1)}{2} \right]}$, $\langle r \rangle_{n, l_{\max}} = a_0 \left(n^2 + \frac{n}{2} \right)$

"Hydrogenic" atoms

Consider an atom with one electron, but a more highly charged nucleus, e.g. singly ionized He, atomic number $Z=2$



The Hamiltonian is the same except the Coulomb potential

$$V(r) = \frac{q_1 q_2}{r} = \frac{Z e^2}{r}$$

The characteristic scales of length and energy now change, but how?

Recall $a_c \equiv \frac{\hbar}{m_e v_c}$ $E_c = \frac{\hbar^2}{a_c^2 m_e} = \frac{q_1 q_2}{a_c} = \frac{Z e^2}{a_c}$

$$\Rightarrow a_c = \frac{\hbar^2}{Z m_e e^2} = \frac{a_0}{Z}$$
$$\Rightarrow E_c = Z^2 \frac{e^2}{a_0} = Z^2 E_0$$

Thus the solutions to the hydrogen atom can now be directly used with this scaling

$$E_n = -Z^2 \frac{R}{n^2}$$

$$\rightarrow$$

More generally, we can consider a "hydrogenic" atom consisting of the bound states of two arbitrary charges with arbitrary masses. Such atoms are called "hydrogenic" because they have the same universal form of hydrogen, but just a different scaling.

Now we must consider the reduced mass

$$E_c = \frac{p_c^2}{\mu} = \frac{\hbar^2}{\mu a_c^2} = \frac{q_1 q_2}{a_c}$$

$$\Rightarrow a_c = \frac{\hbar^2}{\mu q_1 q_2}$$

$$\Rightarrow E_c = \frac{\mu (q_1 q_2)^2}{\hbar^2}$$

Energy levels of bound states: $E_n = -\frac{E_c}{2n^2}$

Energy eigenfunctions $\psi_{n,\ell,m} = R_{n,\ell}(r) Y_\ell^m(\theta, \phi)$

$$R_{n,\ell}(r) = N_{n,\ell} \left(\frac{r}{na_c}\right)^\ell e^{-\frac{r}{na_c}} \sum_{l=n-\ell-1}^{2\ell+1} \left(\frac{2r}{na_c}\right)$$